

THEORY

1. The reaction of a metal with an acid produces a salt and hydrogen gas. For example, zinc reacts with hydrochloric acid to form zinc chloride and hydrogen gas.

$$\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$$

2. The reaction of a metal with an acid is an exothermic reaction. It releases heat energy.

3. The reaction of a metal with an acid is a redox reaction. The metal is oxidized and the acid is reduced.

Observation	Chemical Reaction
Effervescence is observed. The gas evolved is hydrogen gas.	$\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
The reaction is exothermic. Heat is evolved.	
The metal is oxidized and the acid is reduced.	

QUESTION

