

THEORY

When a gas is heated, its volume increases. This is because the gas molecules gain kinetic energy and move faster, pushing against the walls of the container. The pressure exerted by the gas molecules increases, causing the volume to expand. This relationship is described by Charles's Law, which states that the volume of a gas is directly proportional to its temperature when the pressure is constant.

Temperature (°C)	Volume (ml)
0	100
10	105
20	110
30	115
40	120
50	125
60	130
70	135
80	140
90	145
100	150

EXPERIMENT



Temperature (°C)	Volume (ml)
0	100
10	105
20	110
30	115
40	120
50	125
60	130
70	135
80	140
90	145
100	150