

EX. 11.11.10. COPPER AND ZINC PH. CELL

Answer: (a)

(b) (i) $Zn \rightarrow Zn^{2+} + 2e^{-}$

(ii) $Cu^{2+} + 2e^{-} \rightarrow Cu$

(iii) $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$

(iv) $E_{cell} = 1.10V$

(v) $Q = 2 \times 96500 = 193000C$

(vi) $n = 2$

(vii) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(viii) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(ix) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(x) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xi) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xii) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xiii) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xiv) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xv) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

(xvi) $Q = nF$

$193000 = 2 \times F$

$F = 96500C$

