

THEORY

The theory of the present experiment is based on the fact that the rate of reaction between a metal and an acid is directly proportional to the surface area of the metal. In this experiment, the rate of reaction between zinc and hydrochloric acid is studied. The reaction is as follows:

$$\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$$

The rate of reaction is measured by the volume of hydrogen gas evolved over a fixed period of time. The rate of reaction is expected to increase with the surface area of the zinc metal.

AIM

To study the effect of surface area of zinc metal on the rate of reaction with hydrochloric acid.

Prepared by: _____

EXPERIMENT



Prepared by: _____