

THEORY

The theory of the present experiment is based on the fact that the rate of reaction between a metal and an acid is directly proportional to the surface area of the metal. In other words, the rate of reaction is directly proportional to the length of the metal strip used. This is because the larger the surface area of the metal, the more acid molecules will be in contact with it, leading to a faster reaction rate.

APPARATUS

- 1. Dilute hydrochloric acid
- 2. Magnesium ribbon
- 3. Test tube
- 4. Delivery tube
- 5. Gas jar
- 6. Water trough
- 7. Stopwatch
- 8. Measuring cylinder
- 9. Glass plate
- 10. Stand and clamp

PROCEDURE

1. A small amount of dilute hydrochloric acid is poured into a test tube.

EXPERIMENT

