

EXAMINATION

1. The following information is given for the reaction:

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$

Calculate the mass of water formed when 4g of hydrogen reacts with 32g of oxygen.

| Element | Relative Atomic Mass | Relative Formula Mass |
|----------------------|----------------------|-----------------------|
| H | 1 | |
| O | 16 | |
| H_2 | | 2 |
| O_2 | | 32 |
| H_2O | | 18 |

ANSWER: 36g

KINGSTON



| Relative Atomic Mass | Relative Formula Mass |
|----------------------|-----------------------|
| H = 1 | |
| O = 16 | |
| $\text{H}_2 = 2$ | |
| $\text{O}_2 = 32$ | |

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